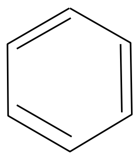
BENZENE

Structure Primary analysis revealed benzene had an... **empirical formula of CH**
and a **molecular formula of C_6H_6**

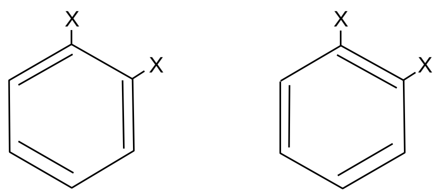
Q.1 *Draw out some suitable structures which fit the molecular formula C_6H_6*

Kekule Kekulé suggested a **PLANAR**, **CYCLIC** structure with **ALTERNATING DOUBLE AND SINGLE BONDS**

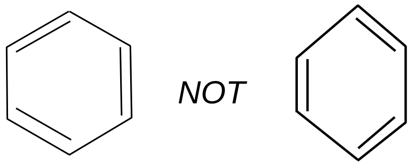


- However it **did not readily undergo electrophilic addition** *- no true C=C bond*

- only **one 1,2 disubstituted product** existed
  i.e you didn't get two isomers like these..

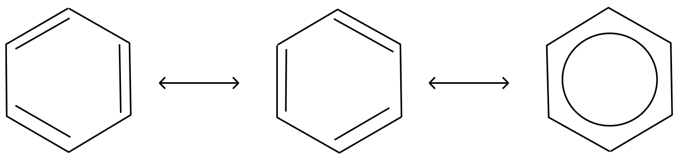


- all six C—C **bond lengths were similar**.
  Double bonds are shorter than single ones



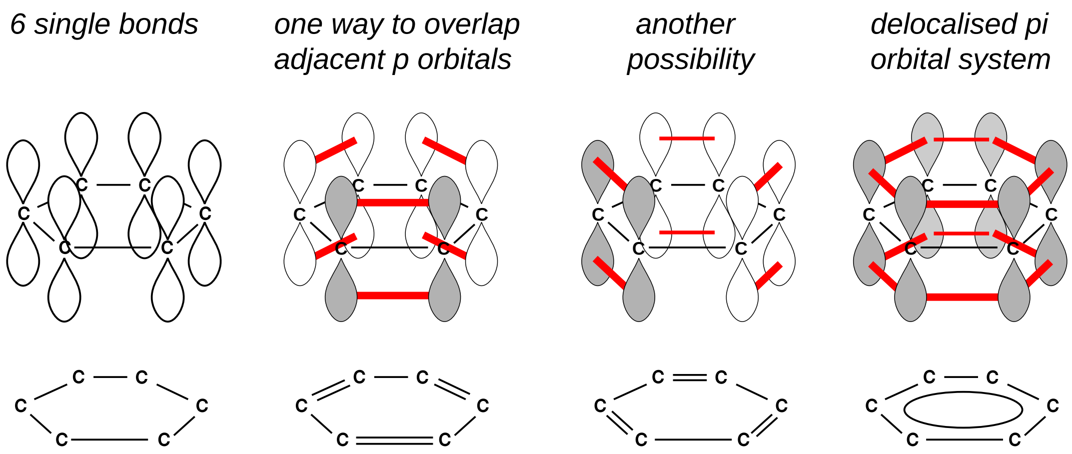
- the ring was **thermodynamically more stable** than expected - *see below*

To explain the above, it was suggested that the structure oscillated between the two Kekulé forms but was represented by neither of them - a **resonance hybrid**.

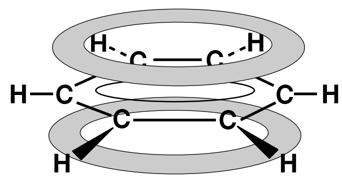


Delocalised system

The theory suggested that instead of three localised (in one position) double bonds, the six π (pi) electrons making up those bonds were **delocalised** (not in any one particular position) around the ring by overlapping the p orbitals. There would be no double bonds to be added to and all bond lengths would be equal. It also gave a **planar** structure.

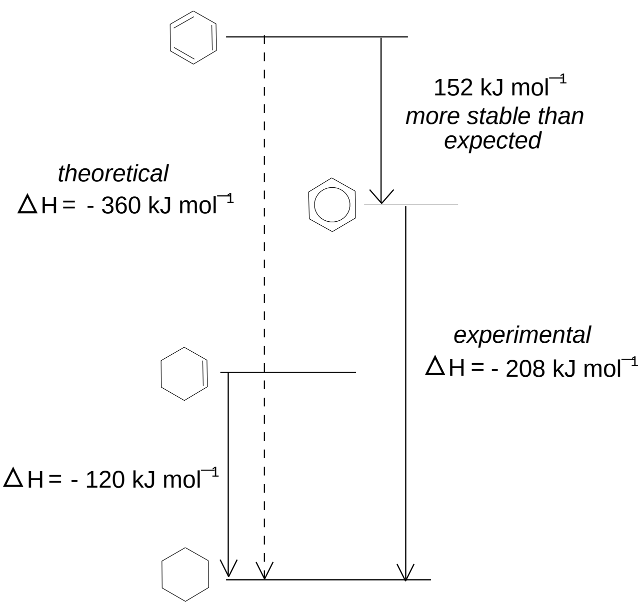


This **structure was particularly stable** and resisted any attempt to break it down through normal electrophilic addition. However, overall substitution of any of the hydrogen atoms would not affect the delocalised system.



Thermodynamic evidence for stability

When unsaturated hydrocarbons are reduced to the corresponding saturated compound, energy is released. If the experiment is carried out in a **bomb calorimeter** the amount of heat liberated per mole (enthalpy of hydrogenation) can be measured.



When cyclohexene (one C=C bond) is reduced to cyclohexane, 120kJ of energy is released per mole.

$$C_6H_{10}(l) + H_2(g) \longrightarrow C_6H_{12}(l)$$

Theoretically, if benzene contained three separate C=C bonds it would release 360kJ per mole when reduced to cyclohexane

$$C_6H_6(l) + 3H_2(g) \longrightarrow C_6H_{12}(l)$$

- benzene releases 208kJ per mole when reduced putting it lower down the energy scale
- it is 152kJ per mole more stable than expected.
- this value is known as the **resonance energy.**

REACTIONS OF THE BENZENE (AROMATIC) RING

Nitration Converts benzene into nitrobenzene, $C_6H_5NO_2$

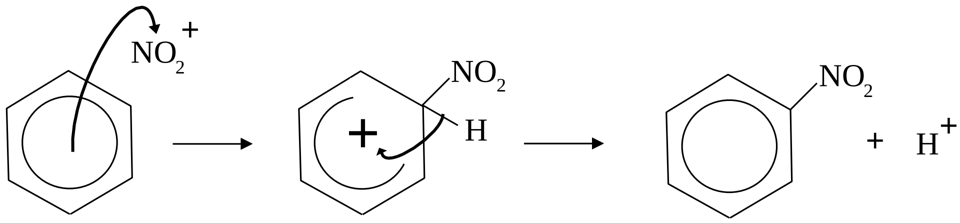
The nitration of benzene is the first step in an historically important chain of reactions. These lead to the formation of dyes, and explosives.

reagents **conc**. nitric acid and **conc**. sulphuric acid (catalyst)

conditions reflux at 55°C

equation $C_6H_6 + HNO_3 \longrightarrow C_6H_5NO_2 + H_2O$

mechanism



- an electron pair leaves the delocalised system to form a bond to the electrophile
- this disrupts the stable delocalised system and forms an unstable intermediate.
- to restore stability, the pair of electrons in the C-H bond moves back into the ring.
- overall there is substitution of hydrogen ... **ELECTROPHILIC SUBSTITUTION**

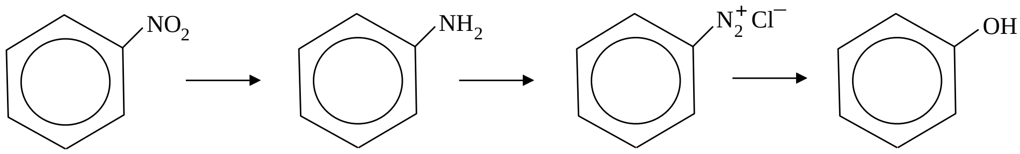
electrophile NO_2^+ , **nitronium ion** or nitryl cation

it is generated in an **acid-base** reaction as follows...

$2H_2SO_4 + HNO_3 \rightleftharpoons 2HSO_4^- + H_3O^+ + NO_2^+$

$2H_2SO_4$	HNO_3
acid	*base*
proton donor	*proton acceptor*

Importance Nitrobenzene is the start of an important chain of reactions.



nitrobenzene — phenylamine (aniline) — benzene diazonium chloride — phenol

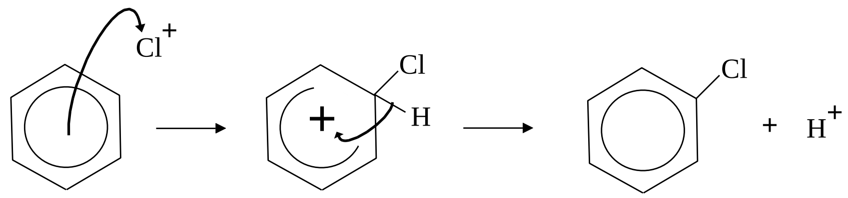
Halogenation Converts benzene into chlorobenzene, C_6H_5Cl

reagents chlorine and a halogen carrier (the catalyst)

conditions reflux in the presence of a **halogen carrier** such as **iron**, **iron(III) chloride**, **iron(III) bromide**, **aluminium chloride**

equation $C_6H_6 + Cl_2 \longrightarrow C_6H_5Cl + HCl$

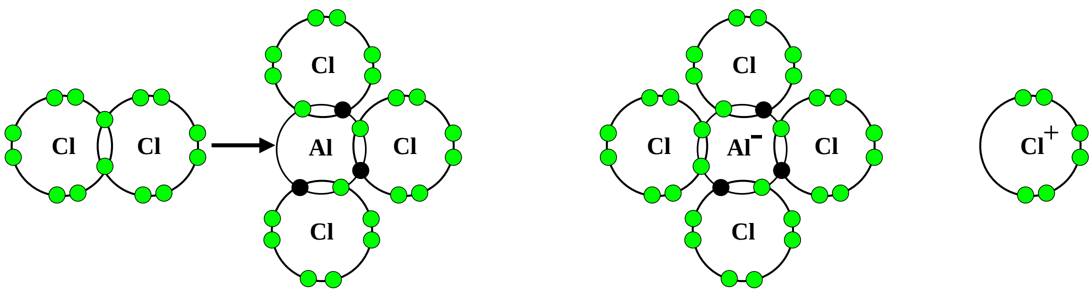
mechanism



PROBLEM *Chlorine is non polar so is not a good electrophile. A catalyst (HALOGEN CARRIER) is required to polarise the halogen.*

- **anhydrous aluminium chloride** can act as the catalyst
- the Al in $AlCl_3$ has only 6 electrons in its outer shell; - **LEWIS ACID**
- it **increases the polarisation of the Cl-Cl bond**
- this makes the charge on C more positive and the following occurs

$$Cl_2 + AlCl_3 \rightleftharpoons AlCl_4^- + Cl^+$$



The aluminium chloride acts as a **LEWIS ACID** as the aluminium atom has only 6 electrons in its outer shell. It accepts a pair of electrons from one of the atoms in the chlorine molecule.

All the atoms in the new species have complete outer shells

The new species has 6 electrons in its outer shell. It has one less electron than a chlorine atom so has a **+** charge.

- the H^+ ion removed from the benzene ring reacts with the $AlCl_4^-$ ion to regenerate the $AlCl_3$

$$H^+ + AlCl_4^- \longrightarrow AlCl_3 + HCl$$

Comparison with alkenes The conditions are much tougher than with alkenes because the delocalised system makes benzene more stable and less reactive.